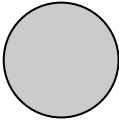






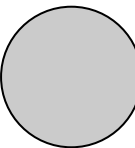
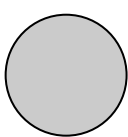
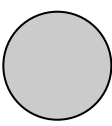
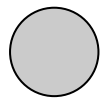
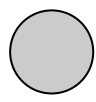
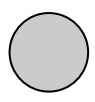
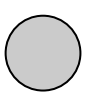
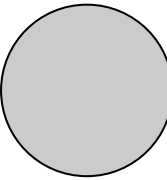
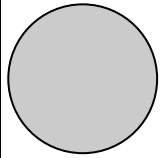
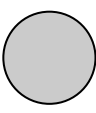
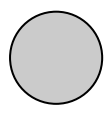
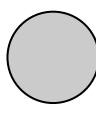
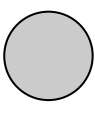
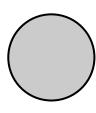
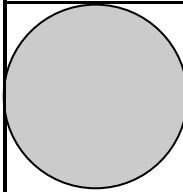
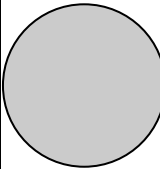
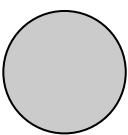
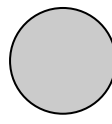
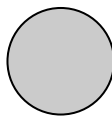
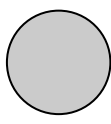
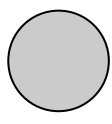




Atomradius und Periodensystem

Maßeinheit Picometer: $1 \text{ pm} = 10^{-12} \text{ m} = \frac{1}{1.000.000.000.000} \text{ m}$

I	II	III	IV	V	VI	VII
 152 Li Lithium	 _____ Be Beryllium	 _____ B Bor	 _____ _____ Kohlenstoff	 _____ N Stickstoff	 _____ O Sauerstoff	 64 F Fluor
 _____ Na Natrium	 160 Mg Magnesium	 143 _____ Aluminium	 117 Si Silicium	 110 P Phosphor	 104 S Schwefel	 99 Cl Chlor
 _____ K Kalium	 _____ Ca Calcium	 122 Ga Gallium	 122 Ge Germanium	 121 As Arsen	 117 Se Selen	 _____ Br Brom
 _____ Rb Rubidium	 215 Sr Strontium	 162 In Indium	 140 Sn Zinn	 141 Sb Antimon	 137 Te Tellur	 _____ I Iod

Innerhalb einer Hauptgruppe nimmt der Atomradius von _____ nach _____ zu. Innerhalb einer _____ nimmt der Atomradius bei gleichbleibender Außenschale von links nach rechts _____.

Arbeitsauftrag:

- Ergänze mit Hilfe des CHEMIE-MASTER® - Periodensystems die fehlenden Radien und Symbole.
- Begründe schriftlich:
 - Warum ist das Fluor-Atom kleiner als das Sauerstoff-Atom, obwohl beide gleich viele Schalen besitzen?
 - Erkläre, warum Natrium und Kalium unterschiedliche Atomradien haben.